

# CHEM2006A: PRINCIPLES OF INORGANIC CHEMISTRY

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## Effective Term

Semester A 2022/23

## Part I Course Overview

### Course Title

Principles of Inorganic Chemistry

### Subject Code

CHEM - Chemistry

### Course Number

2006A

### Academic Unit

Chemistry (CHEM)

### College/School

College of Science (SI)

### Course Duration

One Semester

### Credit Units

3

### Level

B1, B2, B3, B4 - Bachelor's Degree

### Medium of Instruction

English

### Medium of Assessment

English

### Prerequisites

Nil

### Precursors

Nil

### Equivalent Courses

BCH2006A Principles of Inorganic Chemistry

### Exclusive Courses

Nil

### Additional Information

Note: CHEM2006A does not contain any practical component, and has a credit unit value of three (3).

## Part II Course Details

### Abstract

The aim of this course is to allow students to develop an understanding of the basic principles and modern concepts of inorganic chemistry from a discovery approach with an emphasis on atomic structures and properties, chemical bonding and coordination chemistry; Students undertaking BCH2006 will also acquire practical experience on inorganic chemistry experiments.

### Course Intended Learning Outcomes (CILOs)

	CILOs	Weighting (if app.)	DEC-A1	DEC-A2	DEC-A3
1	Relate the atomic structures of elements with their physical and chemical properties.	20	x		
2	Apply Lewis structures, Valence-Shell Electron-Pair Repulsion Theory, Valence Bond Theory and Molecular Orbital Theory to explain the physical and chemical properties of molecules, leading to the design and discovery of new molecules.	20	x	x	
3	Describe and explain the chemistry of Groups 17 and 18 elements.	10	x	x	
4	Describe possible isomerism of a metal complex and explore the role of the electronic configuration of a metal complex on its coordination number, geometry and reactivity.	15	x		
5	Predict and explain the colour, stability, geometry and magnetic properties of a metal complex using Valence Bond Theory and Crystal Field Theory.	20	x	x	
6	Explain the properties of solvents, acids and bases.	15	x		
7	Apply principles and synthetic techniques to conduct inorganic chemistry experiments from a discovery approach.	0		x	x

#### A1: Attitude

Develop an attitude of discovery/innovation/creativity, as demonstrated by students possessing a strong sense of curiosity, asking questions actively, challenging assumptions or engaging in inquiry together with teachers.

#### A2: Ability

Develop the ability/skill needed to discover/innovate/create, as demonstrated by students possessing critical thinking skills to assess ideas, acquiring research skills, synthesizing knowledge across disciplines or applying academic knowledge to real-life problems.

#### A3: Accomplishments

Demonstrate accomplishment of discovery/innovation/creativity through producing /constructing creative works/new artefacts, effective solutions to real-life problems or new processes.

## Teaching and Learning Activities (TLAs)

	TLAs	Brief Description	CILO No.	Hours/week (if applicable)
1	Lectures and tutorials	Students will develop an understanding on atomic structures and correlation with the physical and chemical properties in lectures and tutorials from a discovery approach.	1	0.5
2	Lectures and tutorials	In lectures and tutorials, students will apply various theories to explore, predict and explain the structures, bonding and properties of simple molecules, leading to the design and discovery of new molecules.	2	1
3	Lectures and tutorials	In lectures and tutorials, students will develop an understanding on the chemistry of Groups 17 and 18 elements	3	0.5
4	Lectures and tutorials	Students will develop an understanding of isomerism of a metal complex and correlate properties of a metal complex with its electronic configuration in lectures and tutorials.	4	0.5
5	Lectures and tutorials	In lectures and tutorials, students will apply Valence Bond Theory and Crystal Field Theory to explain various properties of a metal complex from a discovery approach.	5	0.5
6	Lectures and tutorials	Students will develop an understanding of solvents, acids and bases and predict their properties in lectures and tutorials.	6	0.5

7	Experiments and written reports	Students, in the form of small groups, will conduct inorganic chemistry experiments from a discovery approach, and present their results in written reports. (CHEM2006 only)	7	2
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**Assessment Tasks / Activities (ATs)**

ATs	CILO No.	Weighting (%)	Remarks (e.g. Parameter for GenAI use)
1	Tutorial Questions and Assignments	1, 2, 3, 4, 5, 6	15
2	Tests	1, 2, 3, 4, 5, 6	15

**Continuous Assessment (%)**

30

**Examination (%)**

70

**Examination Duration (Hours)**

3

**Additional Information for ATs**

Starting from Semester A, 2015-16, students must satisfy the following minimum passing requirement for courses offered by CHEM:

“A minimum of 40% in both coursework and examination components.”

**Assessment Rubrics (AR)****Assessment Task**

Tutorial questions and assignments

**Criterion**

ABILITY to develop an understanding on basic concepts of inorganic chemistry

**Excellent (A+, A, A-)**

High

**Good (B+, B, B-)**

Significant

**Fair (C+, C, C-)**

Moderate

**Marginal (D)**

Basic

**Failure (F)**

Not even reaching marginal levels

**Assessment Task**

Tests

**Criterion**

ABILITY to describe and explain basic concepts of inorganic chemistry to solve problems

**Excellent (A+, A, A-)**

High

**Good (B+, B, B-)**

Significant

**Fair (C+, C, C-)**

Moderate

**Marginal (D)**

Basic

**Failure (F)**

Not even reaching marginal levels

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**Assessment Task**

Lab reports(CHEM2006 only)

**Criterion**

ABILITY to conduct inorganic chemistry experiments and present results in written reports

**Excellent (A+, A, A-)**

High

**Good (B+, B, B-)**

Significant

**Fair (C+, C, C-)**

Moderate

**Marginal (D)**

Basic

**Failure (F)**

Not even reaching marginal levels

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**Assessment Task**

Examination

**Criterion**

ABILITY to describe, explain, and integrate basic concepts of inorganic chemistry and apply them to solve problems

**Excellent (A+, A, A-)**

High

**Good (B+, B, B-)**

Significant

**Fair (C+, C, C-)**

Moderate

**Marginal (D)**

Basic

**Failure (F)**

Not even reaching marginal levels

## Part III Other Information

**Keyword Syllabus**Atomic Structure

The quantisation of energy. Bohr's theory of the hydrogen atom. Quantum theory and atomic orbitals. Many-electron atoms.

Atomic Properties

Atomic radii. Ionisation energies. Electron affinities. Electronegativity. Periodic trends.

Chemical Bonding

Lewis structures. Resonance. VSEPR method and molecular geometry. Hybrid orbitals. Molecular orbitals. Bonding in solids.

Main Group Chemistry

Group 17 chemistry: hydrogen halides, interhalogen compounds, and oxidation state. Group 18 chemistry: VSEPR.

Coordination Chemistry

Coordination numbers and coordination geometries. Types of ligands. Bonding. Reactivity of coordination compounds.

Solvents, Solutions, Acids and Bases

Solvent properties. Definitions of acids and bases. Hard and soft acid base concepts.

**Reading List****Compulsory Readings**

Title	
1	Nil

**Additional Readings**

Title	
1	Inorganic Chemistry 6th Edition, Duward Shriver, Mark Weller, Tina Overton, Fraser Armstrong, Jonathan Rourke, Publisher: W. H. Freeman, 2014.
2	Basic Inorganic Chemistry, F. A. Cotton, G. Wilkinson, P. L. Gaus, 2nd Edition, Wiley, New York, 1995.